Honors Chemistry
Double Replacement Reactions

What is a distinguishing feature of every double replacement reaction?

Separate the following compounds into the ions (type, charge, and number) that are present in an aqueous solution.

Example: $Al_2(SO_4)_3 \Rightarrow 2\ Al^{3+}$ and $3\ SO_4^{2-}$

NaCl _____ and _____

MgF_2 _____ and _____

$AlBr_3$ _____ and _____

K_2S _____ and _____

$AgNO_3$ _____ and _____

$CuSO_4$ _____ and _____

Li_3PO_4 _____ and _____

$Ba(OH)_2$ _____ and _____

$Pb(NO_3)_2$ _____ and _____

$KMnO_4$ _____ and _____

HCl _____ and _____

$Fe(NO_3)_2$ _____ and _____

$Fe(HCO_3)_3$ _____ and _____

$Zn(C_2H_3O_2)_2$ _____ and _____

Predict the results of the following double replacement reactions; then <u>balance</u> each of the following equations. *Assume all reactants are aqueous solutions*. <u>Be sure to indicate the physical states of both products.</u>

1. ____ NaCl + ____ $Pb(NO_3)_2 \rightarrow$

2. ____ $AgNO_3$ + ____ HCl $\rightarrow$

3. ____ $Ba(OH)_2$ + ____ $Na_3PO_4 \rightarrow$

4. ____ K_2S + ____ $AgNO_3 \rightarrow$

5. ____ $CaCl_2$ + ____$Na_2CO_3 \rightarrow$

6. ____ FeI_2 + ____LiF $\rightarrow$

7. ____ $NiSO_4$ + ____NaOH $\rightarrow$

8. ____ $Al_2(SO_4)_3$ + ____KF $\rightarrow$